

University of Rajasthan

B.Sc Semester III Chemistry practical Exam (A)

Max Marks-10*40.

Min marks 4*+16

duration _4hrs

Ex1 Estimate Nickel as Nickel dimethyl glyoximate. (10)

EX2(a) Identification of organic compound(solid for liquid) through element detection (N&S) melting point boiling point, functional group analysis with preparation of suitable derivative. (7)

(A-Benzoic acid,(B)aniline)

(b) Synthesis of acetanilide from aniline. (3)

Ex3 To determine the heat of neutralization of given acid and base. (10)

Or

Determination of the velocity constant and order of reaction for the hydrolysis of ethyl acetate by sodium hydroxide at room temperature (saponification of an ester)

Ex4 viva-voice. (5)

Ex5 Record. (5)

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B.Sc Semester III Chemistry practical Exam (B)

Max Marks-10*40.

Min marks 4*+16

Duration _4hrs

Ex1 Estimate copper as cuprous thiocyanate. (10)

Ex2(a) Identification of organic compound (solid or liquid) through element detection(N&S) melting /boiling point function a group analysis with the preparation of suitable derivative. (7)

(A-Ethyle alcohol,B- phenol)

(B) Synthesis of meta-nitro aniline from meta-dinitrobenzene. (3)

Ex3 To determine the dissociation energy of given weak acid. (10)

Or

To study the distribution of benzoic acid between benzene and water

Ex4 viva-voice (5)

Ex5 Record (5)

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B.Sc Semester III Chemistry practical Exam (C)

Max Marks-10*40.

Min marks 4*+16

Duration _4hrs

Ex1 Estimate zinc as zinc ammonium phosphate. (10)

EX2(a) identification of organic compound(solid for liquid) through element detection (N&S)melting point boiling point,functional group analysis with preparation of suitable derivative. (7)

(A-acetic acid ,B-phenol)

(b) Synthesis of Aspirin from salicylic acid. (3)

Ex3 To determine the heat of neutralization of given acid and base. (10)

Or

To determine partition coefficient iodine between water and CCl₄,CHCl₃,CS₂ at room temperature. (10)

Ex4 viva-voice (5)

Ex5 Record (5)

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Solubility in $\text{NaHCO}_3</$

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Add potassium and boil the mixture ~~Result~~.

Result:-

m-nitroaniline is obtained by the above method.

Physical state - Yellow coloured shiny crystals

Yield - 2.5 gm

Melting point - 114°C

Reaction

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